SECONDERVIEUR INSTITUT GÉOPOLYMÈRE

Prof. Dr. Joseph Davidovits

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Geopolymer is a term covering a class of synthetic <u>aluminosilicate</u> materials with potential use in a number of areas, essentially as a replacement for <u>Portland cement</u> and for advanced high-tech composites, ceramic applications or as a form of <u>cast stone</u>.

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The name Geopolymer was first applied to these materials by Joseph Davidovits[1] in the 1970s, although similar materials had been developed in the former <u>Soviet Union</u> since the 1950s, originally under the name "soil cements".[2][3] However, this name never found widespread usage in the English language, as it is more often <u>applied</u> to the description of soils which are consolidated with a small amount of <u>Portland</u> <u>cement</u> to enhance strength and stability.

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Geopolymers are an example of the broader class of <u>alkali</u>activated binders, which also includes alkali-activated <u>metallurgical slags</u> and other related materials.[4]

Much of the drive behind research carried out in academic institutions involves the development of geopolymers as a potential large-scale replacement for <u>concrete</u> produced from Portland cement. This is due to geopolymers' alleged lower <u>carbon dioxide</u> production emissions, greater chemical and thermal resistance and better mechanical properties at both ambient and extreme conditions.

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There is some debate^[by whom?] as to whether geopolymer cement has lower CO₂ emissions compared to Portland cement. Calcination of limestone in production of Portland cement is responsible for CO₂ emissions (one ton of cement produced releases one ton of CO₂), while some processes of formation of lime also release CO₂. Mainly it is the ratio of CO₂ reduction that is under debate, and it is process-dependent.

Production

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Geopolymers are generally formed by reaction of an aluminosilicate powder with an alkaline silicate solution at roughly ambient conditions. <u>Metakaolin</u> is a commonly used starting material for laboratory synthesis of geopolymers, and is generated by thermal activation of <u>kaolinite</u> clay.

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Geopolymers can also be made from sources of <u>pozzolanic</u> materials, such as <u>lava</u> or <u>fly ash[5]</u> from coal. Most studies on geopolymers have been carried out using natural or industrial waste sources of metakaolin and other aluminosilicates.

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Industrial and high-tech applications rely on more expensive and sophisticated siliceous raw materials.

The majority of the Earth's crust is made up of Si-AI compounds. Davidovits proposed in 1978 that a single <u>aluminium</u> and <u>silicon</u>-containing compound, most likely geological in origin, could react in a <u>polymerization</u> process with an alkaline solution. The binders created were termed "geopolymers"

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Some have *alleged* that ancient "Roman cement" is a geopolymer, but in reality this material is chemically unlike alkali activated geopolymers because it is made using lime and forms calcium-silicate-hydrates, making it much closer to Portland cement from a chemical standpoint.

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- 3.Setting via <u>polycondensation</u> reactions into an inorganic "polymer" (actually a crystalline-like lattice).

The inorganic polymer network is in general a highly-coordinated 3dimensional aluminosilicate gel, with the negative charges on tetrahedral Al(III) sites charge-balanced by <u>alkali metal cations</u>.

History

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However, because Roman cement forms calcium-silicate-hydrates, and requires calcined limestone as a reactant/precursor, it is more similar to Portland cement than alkali-activated "geopolymer cements" such as Pyrament cement of LoneStar.[6][7]

References

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- 4. ^ Shi, Caijun; Krivenko, Pavel V.; Roy, Della M. (2006). Alkali-Activated Cements and Concretes. Abingdon, UK: Taylor & Francis.
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...The reaction produces SiO4 and AIO4, tetrahedral frameworks linked by shared oxygens as polysialates or polysialate—siloxo or polysialate—disiloxo depending on the SiO2/AI2O3 ratio in the system. The connection of the tetrahedral frameworks is occurred via long range covalent bonds. Thus, geopolymer structure is perceived as dense amorphous phase comprising of semi-crystalline 3-D alumino-silicate microstructure.

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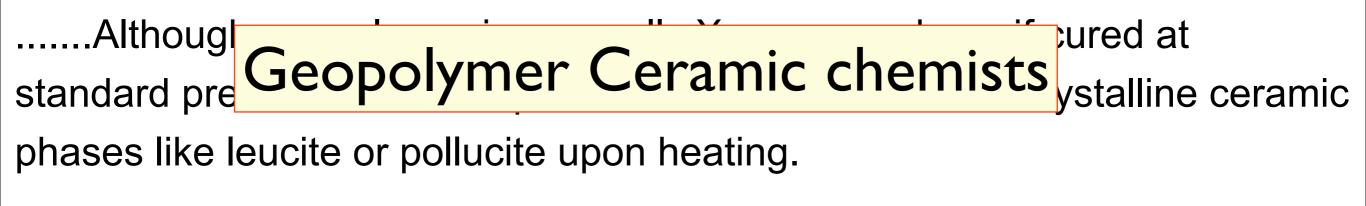
.....Although geopolymer is generally X-ray amorphous if cured at standard pressures and temperatures, it will convert into crystalline ceramic phases like leucite or pollucite upon heating.

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MK-750 alumoxyl group: Si-O-Al=O

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MK-750 alumoxyl group: Si-O-Al=O reacts in both systems.